**Unit 2, Part 1 Completed Review**

1. pH scale-a scale that measures the strength of an acid or base in a solution
2. Physical Change-change in size, shape, or state of matter
3. Chemical Change-change from one substance to another
4. Activation Energy-energy needed to begin the change from potential energy to kinetic energy
5. Element-substance that cannot be broken down to simpler substances by ordinary means
6. Symbol-shorthand way to represent one atom of an element
7. Acids-ionic compounds that break apart in water to form positively charged hydrogen ions
8. Nucleus-positively charged center of an atom containing almost all the atomic mass
9. Proton-positively charged particle in the nucleus, has mass (1 AMU)
10. Electron-negatively charged particle in the electron cloud, no mass
11. Neutron-neutral (no charge) particle in the nucleus, has mass (1 AMU)
12. Atomic Number- the number of protons (or electrons) in one atom of an element
13. Atomic Mass Number-the number of protons PLUS the number of neutrons in an atom of an element
14. Solution-mixture in which one or more substances are uniformly distributed in another substance
15. Isotope-atoms of the same element (same atomic #) with different numbers of neutrons (different atomic mass)
16. Energy level-paths that electrons follow around the nucleus
17. Chemically active-elements that will chemically combine
18. Inert-elements with almost no chemical activity due to full outer energy levels
19. Covalent bond-bond formed when two atoms share electrons in their outer energy level, electrons orbit nuclei of both atoms
20. Ionic bond-bond due to electrical attraction of two atoms which have transferred electrons from one to the other
21. Concentration-amount of solute dissolved in a fixed amount of the solution
22. Ion-atoms which carry + or – electrical charge due to loss or gain of electrons
23. Diatomic molecule-compound formed when two atoms of the same element covalently bond
24. Catalyst-chemical substance that can reduce the amount of activation energy needed to start a reaction
25. Chemical bond-force holding elements (atoms) together to form compounds
26. Polar-uneven distribution of charge
27. Cohesion-water molecules stick to each other as a result of attractive forces between hydrogen
28. Adhesion-attractive forces between two particles of different substances, such as water and glass
29. Base-ionic compounds that break apart in water to form negatively charged hydroxide ions
30. Solute-substance being dissolved